

Compounds

There are millions of different **compounds** around you. Chances are everything you can see is one type of **compound** or another. When **elements** join and become **compounds**, they lose their individual traits. Sodium alone is very reactive. But when sodium and chlorine combine, they form a non-reactive substance called sodium chloride (Salt, NaCl). The **compound** has none of the traits or the original **elements**. The new **compound** is not as reactive as the original **elements**. It has a new life of its own.

Compounds are groups of two or more **elements** that are bonded together. There are two main types of bonds that hold those **atoms** together, covalent and electrovalent/ionic bonds. Covalent **compounds** happen when the **atoms** share the electrons, and ionic **compounds** happen when electrons are donated from one **atom** to another.

When we talk about **compounds**, bonds are built and broken down by chemical forces. Physical forces (unless you're inside of the Sun or something extreme) cannot break down **compounds**. Chemical forces are forces caused by other **compounds** or **molecules** that act on substances.

